



## QUEEN'S VALLEY SCHOOL QUEEN'S VALLEY JUNIOR SCHOOL

### ASSIGNMENT

For Class : IX

Sections : All Sections

Subject: CHEMISTRY

- Q1. Two atoms of an element have atomic number 17 but mass number 35 & 37 respectively, what is the relation between them?
- Q2. Which of the two chemicals would be more reactive – element X of atomic number 18 or element Z of atomic number 16 and why?
- Q3. Atomic number of aluminium is 13 and chlorine is 17. How many electrons, protons and neutrons are present in aluminium and chloride ion?
- Q4. If K & L shells of an atom are completely filled, then what will be the name of the element?
- Q5. How many electrons, protons and neutrons will be there in an element  $X^{19}_9$ . What will be the valency of this element?
- Q6. In the following table are given mass number and atomic number

Element	A	B	C	D	E	F	G	H
Mass No.	2	3	3	6	9	11	19	23
Atomic No.	1	1	2	3	4	5	9	11

- (i) How many neutrons are present in F?
- (ii) Which atoms are isotopes of same element?
- (iii) Which atom will form singly positive charged ion?
- (iv) Which is atom of inert gas?
- (v) Which will form singly negatively charged ion?
- (vi) Which of them has eleven electrons?
- Q7. An atom has electronic configuration 2,8,7.
- (i) What is atomic number of element?
- (ii) Which of the following elements have same number of valence electrons?  
N (7), F(9), P(15), Ar(18)
- Q8. (i) Name the scientists who first proposed a model for the structure of atom.
- (ii) What is maximum number of electron that be accommodated in the outermost orbit of an atom?
- (iii) Draw the atomic structure of sulphur atom showing the arrangement of all its 16 electrons.
- (iv) Define valency by taking example of magnesium.
- Q9. State the new concepts incorporated by Neils Bohr in his proposed model of atom. Draw a diagram to illustrate this model.